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1 **Beyond γ -Al₂O₃ Crystallite Surfaces: the Hidden Features of Edges Revealed by Solid-**
2 **State ¹H NMR and DFT Calculations**

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10

11 **Abstract**

12 Elucidating the nature of high surface area gamma alumina sites is of great interest for
13 numerous applications. In this work, the structural and spectroscopic features of edge sites are
14 unravelled thanks to density functional theory (DFT) calculations combined with high field
15 ¹H MAS NMR of two high surface area alumina samples of distinct morphologies. DFT
16 chemical shift calculations were carried out for relevant surface models with different
17 hydration degrees. However, the best assignment is achieved by considering the first DFT
18 model representing the hydroxylated edges located at the intersection of (110) and (100)
19 alumina facets. The sharp ¹H NMR peak at 0 ppm corresponds to μ_1 -OH groups which are
20 located on this edge and are free from hydrogen bonding. Moreover, we show that these edge
21 sites are the most reactive with respect to chlorine exchange.

22

23 **Keywords:** alumina, density functional theory, edge, hydroxyls, NMR spectroscopy

24

25

26 **1. Introduction**

27 The gamma polymorph of Al_2O_3 is used in numerous industrial applications thanks to its
28 remarkable catalytic and adsorptive properties[1] which have been extensively studied by
29 experimental and theoretical approaches.[2–9] In particular, surface hydroxyl groups,
30 responsible for Brønsted acidity, have been characterized by infra-red (IR) spectroscopy
31 [2,3,10–12] and by ^1H solid-state nuclear magnetic resonance (NMR) spectroscopy.[12–16]
32 Moreover, DFT calculations[5] enabled the refinement of the empirical assignment of the
33 main IR bands thanks to the determination of the hydration of the three main exposed $\gamma\text{-Al}_2\text{O}_3$
34 surfaces:[17] (110), (100) and (111). In ^1H NMR work, the use of high fields and fast magic
35 angle spinning (MAS) provided improved spectral resolution, revealing various, partially
36 resolved, ^1H signals. Three main spectral regions were identified at around 0 ppm, 1-3 ppm
37 and 3-5 ppm, respectively assigned to non H-bonded $\mu_1\text{-OH}$; $\mu_2\text{-OH}$ and $\mu_3\text{-OH}$. These
38 hydroxyls are connected to one, two or three Al atoms exhibiting different coordination
39 (Al_{IV} , Al_{IV} , Al_{VI}).[15,16] Also, broad signals above 5 ppm were associated with hydrogen
40 bond donor species.[16]

41 However, to go beyond the current knowledge of high surface aluminas, atomic scale insights
42 into the nature and location of the hydroxyls that originate each signal are required. In
43 particular, considering extended surfaces only to interpret NMR data overlooks that $\gamma\text{-Al}_2\text{O}_3$
44 crystallites are finite nano-objects exhibiting various morphologies. Like metallic nano-
45 particles with stepped surfaces and edges that provide low coordination sites active in
46 catalysis,[18–21] edge architectures present on alumina crystallites should harbour original
47 hydroxyl and Al sites distinct from those on such surfaces. In a recent review, Busca points to
48 the likely role of edges and corners on alumina's reactivity, which are suspected to be the
49 location of the strongest Lewis acid sites bearing hydroxyl groups more resistant to
50 dehydration.[8,22]

51 Here, the interpretation of the proton NMR spectra of γ -Al₂O₃ is revisited, providing new
52 insights into the nature and location of the hydroxyls in order to refine our current structural
53 knowledge of high surface area aluminas. ¹H NMR spectra was recorded on two Al₂O₃
54 samples of different morphologies, and it is shown by chemical shift DFT calculations that
55 considering alumina facets only leads to an incomplete description of the spectra. This
56 challenge is solved by establishing edge models highlighting the selective chlorination of the
57 edge sites.

58

59 **2. Materials and methods**

60 Two relevant commercial high surface alumina samples are considered, PuralSB3 and TH100
61 (Sasol), labelled *P-egg* and *T-flat* respectively, exhibiting different high BET surfaces
62 (S.I.S1). Their characterization by high-resolution transmission electron microscopy (HR-
63 TEM) suggests that T-flat crystallites are larger and have a better defined and more
64 parallelepipedic platelet-like morphology than P-egg's, which appear to have a round shape
65 (S.I.S1). Chlorinated alumina samples were prepared by exposing alumina to a HCl solution
66 (3.5%_wCl/g_{dry} alumina) for 45 minutes, followed by drying and calcination at 520°C. After
67 thermal treatment under H₂ for 2h at 500°C (ramp 5°C/min) and rotor packing under inert
68 atmosphere, quantitative solid-state ¹H NMR spectra were acquired. H SS MAS NMR spectra
69 were obtained on a SB Bruker Avance III 800 (800 MHz ¹H resonance frequency, 18.8 T)
70 spectrometer using a zirconia 2.5 mm rotor at 30 kHz MAS. Quantitative ¹H spectra were
71 recorded using a DEPTH[23–25] sequence for probe background suppression. Pre-scan delays
72 were set to five times the ¹H longitudinal relaxation time (T₁). Chemical shifts were
73 referenced relative to adamantane. Spectral deconvolution was done using DMFit[26]. These
74 spectra recorded at high magnetic field and relatively fast MAS reveal the surface OH signals
75 and subtle changes in chemical shifts between the two different types of alumina (Figure 1 a).

76 To help for the interpretation of NMR experiments, chemical shifts (S.I.S2.3.) have been
77 calculated by using the linear response approach[27,28] within the PBE-dDsC exchange
78 correlation functional[29,30] and PAW pseudopotentials[31] with an energy cut off of 400 eV
79 as implemented in the VASP code.[32,33] The average of the isotropic chemical shielding (σ)
80 of each proton on a TMS (tetramethylsilane) model (a single molecule surrounded by
81 vacuum) was used as reference to calculate the isotropic chemical shift (δ) of the protons of
82 the various hydroxyls of the alumina surfaces or at the edges :

$$83 \quad \delta_{\text{iso}} = \sigma_{\text{iso}} - \sigma_{\text{ref}} \quad (1)$$

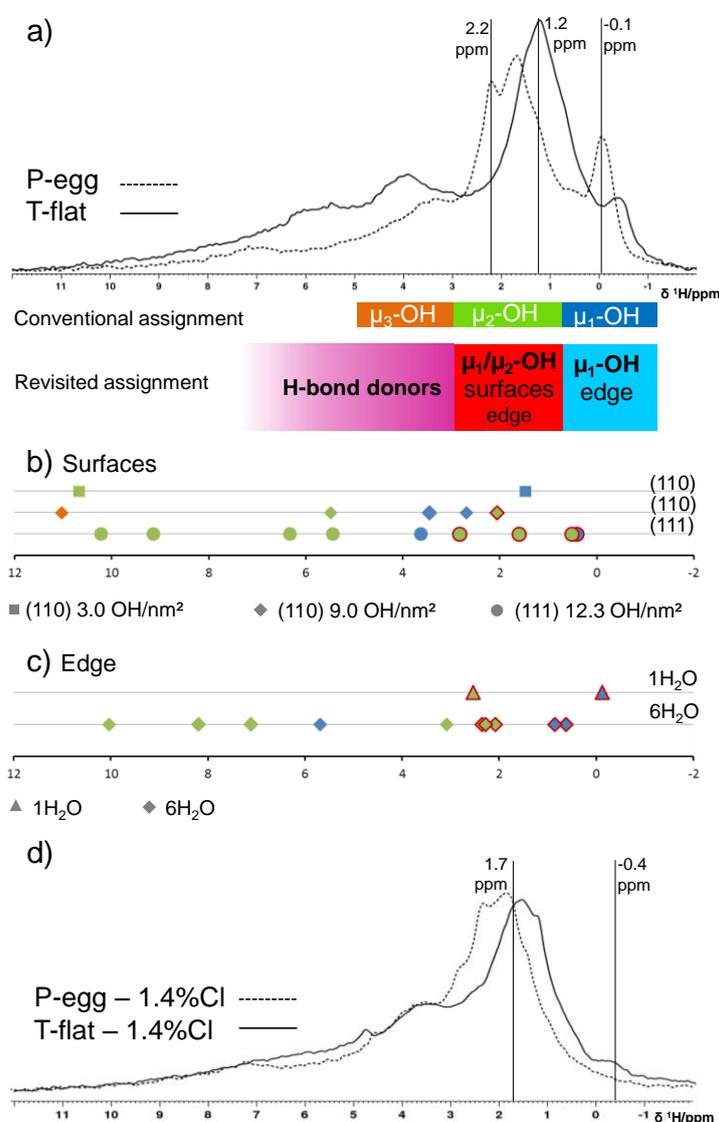
84 For that purpose, relevant periodic models of alumina (110), (100) and (111) surfaces and
85 (110)-(100) edges have been constructed and optimized for various thermodynamically
86 relevant hydration coverages depending on the experimental (T, P) analytical conditions
87 (S.I.S2.). Geometry optimizations were performed using a conjugate-gradient algorithm and
88 convergence criterion on forces of $0.01 \text{ eV } \text{\AA}^{-1}$.

89

90 **3. Results and Discussion**

91 ^1H NMR spectra of P-egg and of T-flat (Figure 1 a) present the main features expected for γ -
92 Al_2O_3 : a resolved signal at around 0 ppm; intense and well-defined peaks in the 1-3 ppm
93 range and broad components ranging from 3-7 ppm. Moreover, the spectra of the two
94 aluminas are clearly distinguished. The most remarkable difference is that the higher-field
95 signal (at respectively -0.1 and -0.4 ppm) is much more intense for P-egg than for T-flat (12%
96 vs 4% of total ^1H signal from spectral deconvolution, Table S2). The 1-3 ppm region is also
97 distinguishable: for P-egg two main signals (1.6 and 2.2 ppm) are observed while for T-flat
98 only one is (1.2 ppm), albeit some shoulders indicating other contributions. Lastly, the
99 contribution of the broad signals ranging from 3-7 ppm is significantly more intense in the T-
100 flat spectrum.

101 In order to rationalize these experimental results, chemical shifts were calculated (Figure 1 b
 102 and S.I.S2.3) for hydrated surface models of the three main exposed γ -Al₂O₃ surfaces[17]
 103 (110), (100) and (111) as defined in previous DFT works [4,5,7] For each surface, several
 104 hydration degrees were considered (S.I.S2.1.) in order to well represent the alumina samples
 105 after thermal treatment, (Figure S8). In the experimental conditions, $10^{-4} < P(\text{H}_2\text{O}) < 10^{-6}$ bar
 106 and $700\text{K} < T < 800\text{K}$, the (110) surface exhibits 3.0 OH/nm² and 9.0 OH/nm², the (111)
 107 surface 12.3 OH/nm², while the (100) surface is dehydrated.



108
 109 **Figure 1.** a) ¹H MAS NMR spectra of two aluminas with different morphologies, P-egg and
 110 T-flat (800 MHz, 30 kHz MAS). The conventional assignment and the revisited one are

111 represented below. b) ^1H chemical shifts calculated by DFT for hydroxyls, colour coded
112 according to the type of OH group given in the conventional assignment, for three surface
113 models; non H-bond donor hydroxyls are highlighted by red outlining. c) ^1H chemical shifts
114 calculated by DFT for hydroxyls for the (110)-(100) edge model at two hydration degrees. d)
115 ^1H MAS NMR spectra (800 MHz, 30 kHz MAS) of chlorinated samples.

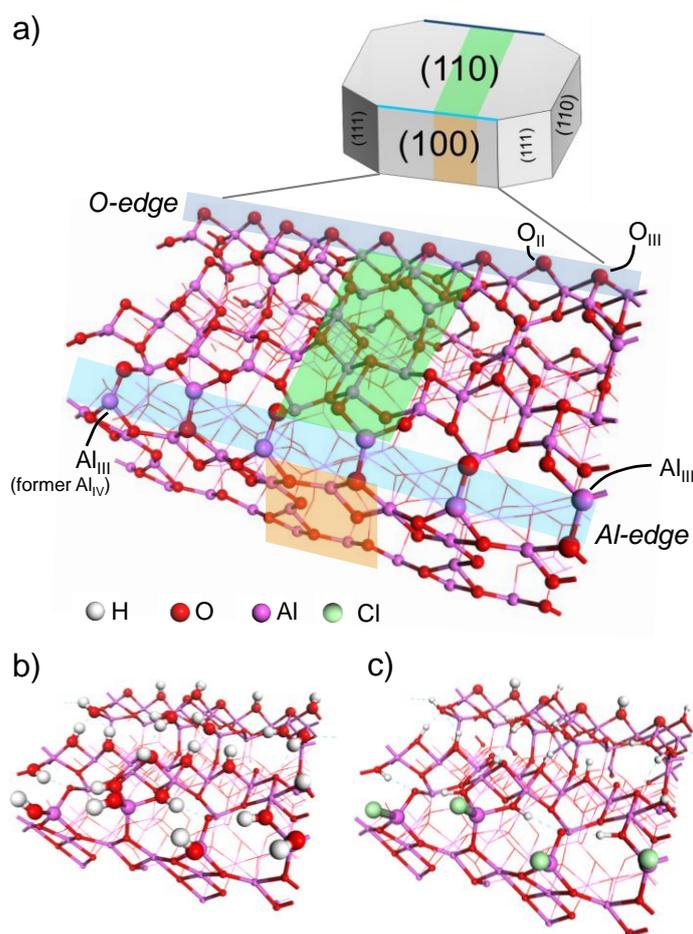
116

117 First, it can be noted that most of the hydroxyls on these surfaces are involved in hydrogen
118 bonding (Tables S8, S9, S10). Almost half of them are H-bond donors, resulting in high
119 chemical shifts: several $\mu_2\text{-OH}$ appearing at >5 ppm. In this case, the chemical shift correlates
120 with hydrogen bond length (Figures S11 and S12),[16,34–36] leading to large variations in δ .
121 Experimentally, H-bond donors are not expected to provide well defined signals such as the
122 ones observed up to 3 ppm.[36]

123 Thus, we observe a poor correspondence between the previously proposed range and the
124 simulated ^1H chemical shifts which do not explain the below 0 ppm signal, expected to
125 correspond to $\mu_1\text{-OH}$. [14,15] Contributions calculated at 0.4 and 0.8 ppm from $\mu_1\text{-OH}$ and $\mu_2\text{-}$
126 OH (respectively) of the (111) surface are found, while the classical assignment does not
127 expect $\mu_2\text{-OH}$ in this range. In addition, no $\mu_3\text{-OH}$ are predicted in the 3-5 ppm region.
128 Overall, these results show that the resonances observed in the experimental ^1H NMR spectra
129 of these $\gamma\text{-Al}_2\text{O}_3$ samples cannot be fully interpreted by considering only their crystalline
130 surfaces. This trend holds true for any hydration degrees (S2.1 and S2.3).

131 To go beyond, another site architecture must be conceived. Apart from the empirical model
132 proposed by Busca,[8] no atomistic model of alumina edge was previously reported in the
133 literature. Thus, a novel model for the (110)-(100) edge was determined (Figure 2 a) based on
134 a nano-rod structure (S.I.S2.2) resulting from the cleavage of the alumina bulk[37] in the two
135 directions perpendicular to the (110) and (100) surfaces. This induces two edge-terminations,

136 one exposing Al atoms (here called Al-edge) and the other exposing O atoms (O-edge). The
 137 Al-edge is constituted of one upper row of Al_{III} atoms (also three-fold coordinated on the
 138 (110) surface) and one lower row of Al_{III} atoms that correspond to Al_{IV} on the (110) surface.
 139 The O-edge exhibits a row of alternating O_{II} and O_{III}, both formally O_{III} on the (110) surface.
 140 These rows of atoms on both edges will be referenced to as edge sites.
 141



142
 143 **Figure 2.** a) Alumina platelet scheme[5] and dehydrated edge model corresponding to two
 144 possible edge terminations (blue) between the (110) (green, top) and the (100) (orange, sides)
 145 surfaces. Edge sites are depicted by bigger balls. b) hydrated edge model with 6H₂O and c)
 146 chlorinated edge model constructed from the 6H₂O hydrated model by exchanging Al-side μ₁-
 147 OH. Blue traced lines indicate hydrogen-bonds (bond length's threshold of 2.5 Å).
 148

149 The systematic study of the hydration state of the nano-rod (and the corresponding edges)
150 shows that water is the most strongly stabilized at the edge sites, rather than on the facet sites.
151 For the adsorption of the first water molecule per unit cell of simulation, this leads to the
152 formation of one μ_2 -OH, the H^+ bonding to the O_{II} atom on the O-edge, and of one μ_1 -OH, the
153 OH^- bonding to the Al_{III} on the lower row on the Al-edge that relaxes into a tetragonal
154 geometry (Table S7). Both these hydroxyls are not involved in hydrogen bonding. The
155 corresponding adsorption energy is $-436 \text{ kJ}\cdot\text{mol}^{-1}$, which is far greater than usual adsorption
156 energies reported on the alumina surfaces,[4] and is consistent with the chemical intuition that
157 the reactivity of Lewis Al edge sites should be greater.[8] Such a configuration is a priori
158 striking when thinking in terms of water dissociation, because the hydroxyl and the
159 corresponding proton are not lying on the same kind of edge. However, one shall consider that
160 experimentally, such low hydration states are obtained upon dehydration of the surface sites,
161 making it possible to leave at the surface distant OH and H groups in the end, after
162 recombination of other OH and H pairs. When more than two water molecules are adsorbed
163 per unit cell of simulation, all the “edge sites” are saturated and the near edge sites on the
164 (110) top surface of the nano-rod start being occupied (Figure 2 b), while the (100) facet of
165 the rod remains dehydrated. These “near edge” sites exhibit a H-bond network identical to
166 that of the surface models. In the conditions of thermal treatment ($10^{-4} < P(H_2O) < 10^{-6}$ bar and
167 $700\text{K} < T < 800\text{K}$), among the multiple hydration degrees equally stable (S.I. Figure S10), we
168 choose two relevant cases: one and six adsorbed water molecules per pair of edges.
169 The calculated 1H chemical shifts for the (110)-(100) edge with one and six adsorbed water
170 molecules are represented in Figure 1 c). μ_1 and μ_2 hydroxyls on edge sites appear in the
171 expected δ range: ≈ 0 ppm and ≈ 2 ppm respectively. These OH are not only free from
172 hydrogen bonding, but also isolated from other neighbouring hydroxyls. For 6 H_2O

173 molecules, hydroxyls with $\delta > 3$ ppm are located on near edge sites and are H-bond donors, as
174 the red outlining indicates.

175 With this edge model a significantly improved correlation between experimental and
176 calculated chemical shifts is achieved, especially for the sharp peak at ≈ 0 ppm. Thus, this
177 peak corresponds mostly to isolated μ_1 -OH located on the edges of alumina crystallites, which
178 are free from hydrogen bonding. While the contribution of some free and H-bond acceptor
179 species on the (111) surface cannot be ruled out, their contribution to the signal is minor, as
180 discussed ahead. The 1-3 ppm region is expected to result from non-isolated μ_1 -OH and μ_2 -
181 OH that are free from H-bonds or H-bond acceptors located on the edges and on the surfaces
182 of the crystallites. The fact that in average, the μ_2 -OH sites are much more represented in this
183 region with respect to the 0-1 ppm region, is in agreement with ^1H - ^{27}Al RESPDOR
184 experiments by Taoufik et al..[15] While not much insight was gained into the 3-5 ppm broad
185 signal, its empirical assignment to μ_3 -OH is strongly questioned. Indeed, for the hydration
186 degree of our samples, only two μ_3 -OH were found in the models of interest (one for (110)
187 12.0 OH/nm^2 and other for edge $6\text{H}_2\text{O}$) and both are hydrogen bond donors with $\delta > 5$ ppm
188 (7.9 and 14.2 ppm, respectively). Moreover, μ_1 -OH and μ_2 -OH species acting as hydrogen
189 bond donors are also impacting the 3-5 ppm broad signal. Lastly, broad signals with $\delta > 5$
190 ppm are thought to correspond to hydroxyls involved in the hydrogen bond network of the
191 surfaces as hydrogen bond donors.

192 With this improved assignment, it is now possible to rationalize the impact the alumina nano-
193 crystallite morphology on the ^1H NMR spectra. As mentioned above, P-egg crystallites are
194 rounded and smaller than those of T-flat, which are more parallelepipedic. Thus, P-egg
195 presents a higher edge to surface ratio than T-flat, which explains the relative intensities of the
196 ≈ 0 ppm peaks of edge μ_1 -OH: 12% and 4% of ^1H signal intensity for P-egg and T-flat
197 respectively (Table S2). In the 1-3 ppm region, the δ value of the most intense peak depends

198 on the sample which is a strong indication of different proportions of exposed surface types
199 for each alumina. Indeed, the electron diffraction analysis (Figure S4) suggests that the (111)
200 termination is more exposed in T-flat crystallites than in P-egg's. Finally, the large signal for
201 $\delta > 3$ ppm is more intense for T-flat, which is explained by a H-bond network between
202 hydroxyls that is more developed on the extended surface planes of T-flat than on P-egg.

203 To get further insights into the surface structure of P-egg and T-flat, the reactivity of the
204 alumina hydroxyls were probed with chlorine. Chlorinated aluminas are used in many
205 catalytic processes[38] but Cl can also be used as a probe for μ_1 -OH.[12,14,39] Chlorinated
206 samples were prepared so as to have 0.5 and 1.4%_{w/w}Cl deposited on each alumina (referred
207 as P-egg-x%Cl and T-flat-x%Cl with x=0.5, 1.4). The effect of chlorine on the ¹H NMR
208 spectra is clearly different for both aluminas (Figure 1 d and Figure S6). For P-egg, the 0 ppm
209 signal disappears completely, as previously reported,[12,14] while the rest of the spectrum
210 remains unchanged. For T-flat, not only does the 0 ppm signal not disappear completely, but
211 it is also observed an intensity increase in the 3-5 ppm region, associated to an intensity
212 decrease for $\delta > 5$ ppm (Table S2).

213 According to the DFT OH/Cl exchange energies (S.I. S2.5, Figure 2 c), the edge μ_1 -OH sites
214 are predominantly exchanged with chlorine out of all the considered hydroxyls. The
215 subsequent exchanged hydroxyls would be the μ_1 -OH of the (110) surface and after that only
216 the μ_1 -OH and μ_2 -OH sites of the (111) surface. Since the ≈ 0 ppm peak completely disappears
217 for P-egg, and no impact is observed on the remaining parts of the spectrum, it is believed that
218 only the μ_1 -OH located on the edges are exchanged with chlorine at 1.4%Cl and not those of
219 the surfaces. For T-flat, since the peak at ≈ 0 ppm does not fully disappear while the 3-5 ppm
220 and $\delta > 5$ ppm regions are perturbed, not only the edge μ_1 -OH are exchanged but also surface
221 hydroxyls (most likely on the (110)), disturbing the H-bond network. The signal remaining at
222 ≈ 0 ppm after chlorination should correspond to the (111) surface μ_1 and μ_2 hydroxyls. Just as

223 the electron diffraction analysis, the NMR results also suggest that the (111) surface is
224 relatively more exposed in T-flat crystallites than in P-egg's. This implies that the number of
225 edge μ_1 -OH of T-flat is not sufficient to exchange 1.4% Cl while it is the case on P-egg.

226

227 **4. Conclusion**

228 The construction of the first DFT model of γ -Al₂O₃ crystallite (110)-(100) edge has allowed a
229 refined ¹H NMR peak assignment. It was found that the sharp peak at 0 ppm corresponds
230 mostly to μ_1 -OH located on the edges of the crystallites that are isolated and free from
231 hydrogen bonding. These hydroxyls are also the most favourably exchanged with chlorine,
232 which can be considered as a selective probe of alumina edges. Overcoming the simple
233 empirical assignment, the 1-3 ppm region corresponds to signals from not only μ_2 -OH but
234 also from μ_1 -OH located either on the surfaces or on the edge, and that are either H-bond
235 acceptors or free hydroxyls. Moreover, hydroxyls that are hydrogen bond donors are abundant
236 on the surfaces and contribute to the high chemical shift broad signals. Hopefully, this novel
237 alumina edge model and the improved assignment open new perspectives to further explore
238 the potential of the edge sites present in industrially relevant high surface alumina crystallites.

239

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